

4/pats

10/532854

JC20 Rec'd PCT/PTO

27 APR 2001

NSC 1005

DESCRIPTION

HIGHLY CORROSION-RESISTANT HOT-DIP GALVANIZED STEEL
PRODUCT EXCELLENT IN SURFACE SMOOTHNESS AND FORMABILITY
AND PROCESS FOR PRODUCING SAME

5

Field of the Invention

The present invention relates to a plated steel
sheet and, in more detail, to a highly corrosion-
resistant plated steel product that can be applied to
various applications and, for example, to household
electrical appliances, automobiles and steel sheets for
building materials.

15 Background Art

There are zinc base plated steel sheets that are
often used as plated steel products excellent in
corrosion resistance. The plated steel sheets are used
in various manufacturing industries such as the
automobile, household electrical appliance and building
material industries. Moreover, plated steel products are
used in various other fields such as the plated steel
wire and hot-dip galvanized steel product fields.

In order to improve the corrosion resistance of such
zinc base plated steel products, the present inventors
have proposed a hot-dip Zn-Al-Mg-Si coated steel sheet
(see, e.g., Japanese Patent Publication No. 3,179,446).

Furthermore, in order to improve the corrosion
resistance of zinc base plated steel products, a zinc
base plated steel sheet that is made excellent in age-
based darkening resistance by adding Ti to the hot-dip
Zn-Al coated steel sheet has been proposed (see, e.g.,
Japanese Unexamined Patent Publication (Kokai) No. 5-
125515). However, the problem that the resultant plated
steel sheet shows poor surface smoothness and formability
was not taken into consideration.

Still furthermore, a zinc-base plated-steel sheet,

the appearance of which is made good by adding Ti, B and Si to a hot-dip Zn-Al-Mg coated steel sheet, has been proposed (e.g., see Japanese Unexamined Patent Publication (Kokai) No. 2001-295015). Although Ti and B
5 are added for the purpose of inhibiting the formation and growth of a $Zn_{11}Mg_2$ phase that makes the appearance of the plated steel sheet poor, the proposal neither considers the problem that the addition makes the surface smoothness and formability of the plated steel sheet
10 poor, nor refers to formation of an intermetallic compound.

However, the surface smoothness and formability are not ensured sufficiently for the above plated steel sheets and other disclosed plated sheets.

15 When the solidification rate of the plating layer is adequately ensured during hot-dip galvanizing, the plating layer solidifies before the Al phase grows significantly. As a result, a problem of the surface smoothness does not arise. However, when the
20 solidification rate of the plating layer is small, the Al phases significantly grow first. Protrusions and recesses are then formed on the plating layer surface, and the resultant plated steel sheet has the problem that the surface smoothness and formability become poor.

25

Disclosure of the Invention

An object of the present invention is to solve the above problems, and provide highly corrosion-resistant plated steel products.

30

As a result of intensively carrying out investigations on the development of a plated steel sheet excellent in surface smoothness and formability, the present inventors have discovered that the surface smoothness and formability can be improved by making the
35 plating layer have a metal structure in which one or at least two of the [Mg₂Si phase], [Al phase], [Zn₂Mg phase] and [Zn phase] are present in a mixture in the matrix of

an [Al/Zn/Zn₂Mg ternary eutectic structure], and making one or at least two of the [Al phase], [Zn₂Mg phase] and [Zn phase] contain a Ti-Al base intermetallic compound, and they have thus achieved the present invention. The aspects of the present invention are as described below.

(1) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability, having on the steel product surface a zinc alloy plating layer composed of 4 to 10% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti and the balance of Zn and unavoidable impurities, the plating layer having a metal structure in which one or at least two of the [Al phase], [Zn₂Mg phase] and [Zn phase] are present in a mixture in the matrix of an [Al/Zn/Zn₂Mg ternary eutectic structure], and the plating layer containing a Ti-Al base intermetallic compound in one or at least two of the [Al phase], [Zn₂Mg phase] and [Zn phase].

(2) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability, having on the steel product surface a zinc alloy plating layer composed of 4 to 22% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti, up to 0.5% by mass of Si and the balance of Zn and unavoidable impurities, the plating layer of the plated steel product having a metal structure in which an [Mg₂Si phase], an [Al phase] and a [Zn₂Mg phase] are present in a mixture in the matrix of an [Al/Zn/Zn₂Mg ternary eutectic structure], and the plating layer containing a Ti-Al base intermetallic compound in one or at least two of the [Al phase] and the [Zn₂Mg phase].

(3) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability, having on the steel product surface a zinc alloy plating layer composed of 4 to 22% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti, up to 0.5% by mass of Si and the balance of Zn and unavoidable

impurities, the plating layer of the plated steel product having a metal structure in which an $[\text{Mg}_2\text{Si}]$ phase, an $[\text{Al}]$ phase, a $[\text{Zn}_2\text{Mg}]$ phase and a $[\text{Zn}]$ phase are present in a mixture in the matrix of an $[\text{Al}/\text{Zn}/\text{Zn}_2\text{Mg}]$ ternary eutectic structure], and the plating layer containing a Ti-Al base intermetallic compound in one or at least two of the $[\text{Al}]$ phase, $[\text{Zn}_2\text{Mg}]$ phase and $[\text{Zn}]$ phase].

(4) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability having, on the steel product surface, a zinc alloy plating layer composed of 4 to 22% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti, up to 0.5% by mass of Si and the balance of Zn and unavoidable impurities, the plating layer of the plated steel product having a metal structure in which an $[\text{Mg}_2\text{Si}]$ phase, an $[\text{Al}]$ phase and a $[\text{Zn}]$ phase are present in a mixture in the matrix of an $[\text{Al}/\text{Zn}/\text{Zn}_2\text{Mg}]$ ternary eutectic structure], and the plating layer containing a Ti-Al base intermetallic compound in one or two of the $[\text{Al}]$ phase and $[\text{Zn}]$ phase].

(5) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability, wherein the Ti-Al base intermetallic compound according to any one of (1) to (4) mentioned above is TiAl_3 .

(6) A highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability, wherein the Ti-Al base intermetallic compound according to any one of (1) to (4) mentioned above is $\text{Ti}(\text{Al}_{1-x}\text{Si}_x)_3$ (wherein $x = 0$ to 0.5).

(7) The highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability according to any one of (1) to (6) mentioned above, wherein the Ti-Al base intermetallic compound contained in an $[\text{Al}]$ phase in the plating layer is present in a Zn-Al eutectoid reaction structure in which Zn phases are condensed.

(8) The highly corrosion-resistant hot-dip

galvanized steel product excellent in surface smoothness and formability according to any one of (1) to (7) mentioned above, wherein the size of a dendrite in an [Al phase] in the plating layer is up to 500 μm .

5 (9) A process for producing the highly corrosion-resistant hot-dip galvanized steel product excellent in surface smoothness and formability according to any one of (1) to (8) mentioned above, comprising the step of
10 adding a Ti-Zn base intermetallic compound to a plating bath.

Brief Description of the Drawings

Fig. 1 (a) is a photomicrograph substituted for drawing (magnification: 1,000 x) of a plating layer of
15 the plated steel product of the present invention, and Fig. 1 (b) is a view showing the distribution state of each structure in the photomicrograph.

Fig. 2 (a) is a photomicrograph substituted for drawing (magnification: 3,500 x) that enlarges the [Al"
20 phase] in Fig. 1, and Fig. 2 (b) is a view showing the distribution state of each structure in the photomicrograph.

Best Mode for Carrying out the Invention

25 The hot-dip galvanized steel product according to the present invention is a plated steel sheet having either a plating layer composed of 4 to 10% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti and the balance of Zn and unavoidable impurities, or a
30 plating layer composed of 4 to 22% by mass of Al, 1 to 5% by mass of Mg, up to 0.1% by mass of Ti, up to 0.5% by mass of Si and the balance of Zn and unavoidable impurities, the plating layer of the plated steel sheet having a metal structure in which one or at least two of
35 the [Mg₂Si phase], [Al phase], [Zn₂Mg phase] and [Zn phase] are present in a mixture in the matrix of an [Al/Zn/Zn₂Mg ternary eutectic structure], and the plating

layer containing a Ti-Al base intermetallic compound in one or at least two of the [Al phase], [Zn₂Mg phase] and [Zn phase].

5 The Al content in the Zn-Al-Mg-Ti base plating layer is restricted to 4 to 10% by mass for the following reasons. When the Al content exceeds 10% by mass, the adhesion of the plating layer decreases. The Al content in the plating layer to which Si is not added must therefore be made 10% by mass or less. Moreover, when
10 the Al content is less than 4% by mass, no Al phase crystallizes as primary crystals, and the problem of lowering of smoothness does not arise.

Accordingly, in the hot-dip galvanized steel product of the invention, it is essential to add Si to the
15 plating layer to ensure adhesion of the plating layer particularly when the Al concentration is as high as greater than 10% by mass.

On the other hand, the Al content in the Zn-Al-Mg-Ti-Si base plating layer is restricted to 4 to 22% by
20 mass for the following reasons. When the Al content is less than 4% by mass, no problem about lowering of smoothness arises because no Al phase crystallizes as primary crystals. When the Al content exceeds 22% by mass, the effect of improving the corrosion resistance is
25 saturated.

The Si content is restricted to 0.5% by mass or less (Si content of 0% by mass being excluded) for the following reasons: although Si has the effect of improving the adhesion, the effect is saturated when the
30 content exceeds 0.5% by mass. The Si content is desirably from 0.00001 to 0.5% by mass, and more desirably from 0.0001 to 0.5% by mass.

Addition of Si to a plating layer having an Al content exceeding 10% by mass is essential. However,
35 even for a plating layer having an Al content of up to 10%, because the effect of improving the adhesion of the plating layer is also significant, addition of Si to the

plating layer of a steel product is effective when the steel product is required to have high adhesion of the plating layer, for example, when the steel product is used as a member to be severely worked. Moreover, as a result of adding Si, a $[Mg_2Si]$ phase crystallizes in the solidification structure of the plating layer. Because the $[Mg_2Si]$ phase has the effect of improving the corrosion resistance of a worked portion, it is more desirable to increase an addition amount of Si so that a metal structure in which the $[Mg_2Si]$ phase is present in a mixture in the solidification structure of the plating layer is formed.

The Mg content is restricted to 1 to 5% by mass for the following reasons: when the Mg content is less than 1% by mass, the effect of improving the corrosion resistance is inadequate; when it exceeds 5% by mass, the plating layer is embrittled, and the adhesion thereof decreases.

The Ti content is restricted to 0.1% by mass or less (Ti content of 0% by mass being excluded) for the following reasons. Ti has the effect of crystallizing a Ti-Al base intermetallic compound and improving the surface smoothness and formability. However, when the Ti content exceeds 0.1% by mass, the steel product after plating has a rough surface, and it has a poor appearance. Moreover, when the Ti content exceeds 0.1% by mass, a Ti-Al base intermetallic compound is condensed in the plating layer surface to decrease the surface smoothness and formability. The Ti content is desirably from 0.00001 to 0.1% by mass. The Ti content is more desirably from at least 0.00001% by mass to less than 0.01% by mass.

For the plated steel product according to the present invention, a metal structure containing at least one of the $[Zn]$ phase, $[Al]$ phase, $[Zn_2Mg]$ phase, $[Mg_2Si]$ phase and a Ti-Al base intermetallic compound is formed in the matrix of an $[Al/Zn/Zn_2Mg]$ ternary eutectic

structure] in the plating layer.

The [Al/Zn/Zn₂Mg ternary eutectic structure] herein is a ternary eutectic structure of an Al phase, a Zn phase and an intermetallic compound Zn₂Mg phase. The Al phase forming the ternary eutectic structure corresponds, for example, to an [Al" phase] (Al solid solution dissolving a Zn phase, and containing a small amount of Mg) at high temperature in an Al-Zn-Mg ternary equilibrium state diagram. The Al" phase at high temperature usually appears at room temperature as a fine Al phase and a fine Zn phase in separation. Moreover, the Zn phase in the ternary eutectic structure dissolves a small amount of Al, and further dissolves in some cases a small amount of Mg (Zn solid solution). The Zn₂Mg phase in the ternary eutectic structure is an intermetallic compound phase present near a Zn content of about 84% by weight in a Zn-Mg binary equilibrium state diagram. As long as the state diagram is observed, it is thought that Si and Ti form no solid solution with each phase, or extremely small amounts of Si and Ti form a solid solution therewith even when a solid solution is formed. Because the amounts cannot be definitely distinguished by conventional analysis, the ternary eutectic structure composed of the three phases is represented by an [Al/Zn/Zn₂Mg ternary eutectic structure].

Furthermore, the [Al phase] is a phase that appears to be an island having a distinct boundary in the matrix of the above ternary eutectic structure. The phase corresponds, for example, to an [Al" phase] (Al solid solution dissolving a Zn phase, and containing a small amount of Mg) at high temperature in the Al-Zn-Mg ternary equilibrium state diagram. The Al" phase at high temperature dissolves Zn and Mg with the amounts differing and depending on the concentrations of Al and Mg in the plating bath. The Al" phase at high temperature usually separates into a fine Al phase and a fine Zn phase at room temperature. An island-like shape

observed at room temperature may be taken as ruins of the Al" phase at high temperature. As long as the state diagram is observed, it is thought that Si and Ti do not form a solid solution with the phase, or the amounts are extremely small even when they form a solid solution therewith. However, because conventional analysis cannot definitely determine the amounts, the phase derived from the Al" phase at high temperature and having the ruins of the shape of the Al" phase is termed an [Al phase] in the present invention. The [Al phase] can be definitely distinguished from the Al phase forming the above ternary eutectic structure by microscopic observation.

Furthermore, the [Zn phase] is a phase that appears to be an island having a distinct boundary in the matrix of the above ternary eutectic structure. Actually, the [Zn phase] sometimes dissolves a small amount of Al and further dissolves a small amount of Mg. As far as the state diagram is observed, it is thought that Si and Ti form no solid solution with the phase, or that the amounts are extremely small even when they form a solid solution therewith. In a microscopic observation, the [Zn phase] can be definitely distinguished from the Zn phase forming the above ternary eutectic structure.

Moreover, the [Zn₂Mg phase] is a phase that appears to be an island having distinct boundary in the matrix of the above ternary eutectic structure, and actually dissolves a small amount of Al, sometimes. As far as the state diagram is observed, it is thought that Si and Ti form no solid solution with the phase, or that the amounts of Si and Ti are extremely small even when they form a solid solution therewith. In a microscopic observation, the [Zn₂Mg phase] can be clearly distinguished from the Zn₂Mg phase forming the above ternary eutectic structure.

Furthermore, the [Mg₂Si phase] is a phase that appears to be an island having a distinct boundary in the solidified structure of the plating layer. As far as the

state diagram is observed, it is thought that Zn, Al and Ti form no solid solution with the phase, or that even when they form a solid solution therewith, the amounts are extremely small. In a microscopic observation of the plating layer, the $[\text{Mg}_2\text{Si}]$ phase can be clearly distinguished.

Furthermore, the Ti-Al base intermetallic compound is a phase that appears to be an island having a distinct boundary in the solidified structure of the plating layer. As far as the state diagram is observed, the intermetallic compound is thought to be TiAl_3 . However, because Si is observed when the Ti-Al base intermetallic compound is analyzed in the plating layer to which Si is added, it is thought that the Ti-Al base intermetallic compound in the plating layer is TiAl_3 that dissolves Si or $\text{Ti}(\text{Al}_{1-x}\text{Si}_x)_3$ ($x = 0$ to 0.5) in which Si is substituted for part of Al.

The Ti-Al base intermetallic compound in the hot-dip galvanized steel product of the present invention is characterized in that the intermetallic compound is present in the $[\text{Al}]$ phase, $[\text{Zn}_2\text{Mg}]$ phase and $[\text{Zn}]$ phase. The contained form of the Ti-Al base intermetallic compound is restricted to sites in the $[\text{Al}]$ phase, $[\text{Zn}_2\text{Mg}]$ phase and $[\text{Zn}]$ phase because a Ti-Al base intermetallic compound present in sites other than the above sites cannot improve the surface smoothness and formability. The Ti-Al base intermetallic compound present in sites in the $[\text{Al}]$ phase, $[\text{Zn}_2\text{Mg}]$ phase and $[\text{Zn}]$ phase is thought to improve the surface smoothness and formability for the following reasons: the Ti-Al base intermetallic compound becomes nuclei of the $[\text{Al}]$ phase, $[\text{Zn}_2\text{Mg}]$ phase and $[\text{Zn}]$ phase, promotes crystallization of these crystals, and many fine structures are formed. That is, when the crystals become fine, the recesses and protrusions of the plating layer surface are suppressed. As a result, the surface becomes smooth, and the friction coefficient is decreased during forming due to a decrease in the

deformation resistance of the plating layer. It is thought that the formability of the plated steel product is thus improved.

The effect is significant particularly in the [Al phase]. The plating layer surface is smoothed and the friction coefficient is lowered by adjusting the size of a dendrite of the [Al phase] to 500 μm or less. The size is desirably up to 400 μm , and more desirably up to 300 μm .

As a result of examining metal structures in many plating layers, they have observed intermetallic compounds each having a size of several micrometers in most of the metal structures. Fig. 1 shows one example of the intermetallic compounds present in an [Al phase]. Fig. 1 (a) is a photomicrograph (magnification: 1,000 x) of a plating layer of a plated steel product in the present invention. Fig. 1 (b) is a view showing the distribution state of each structure in the photomicrograph. It is understood from the view that each structure can be definitely specified with the help of a photomicrograph of the plating layer of a plated steel product in the present invention.

In Fig. 1 (a), a Ti-Al base intermetallic compound is observed in a phase corresponding to an [Al" phase] at high temperature in the Al-Zn-Mg ternary equilibrium state diagram. The Al" phase at high temperature usually appears at room temperature as a fine Al phase and a fine Zn phase in separation by a eutectoid reaction taking place at 277°C in the Al-Zn binary equilibrium state diagram. When a hypo-eutectoid reaction takes place herein, the Al" phase crystallized at high temperature starts to precipitate a Zn phase from a ternary eutectic temperature in the Al-Zn-Mg ternary equilibrium state diagram, and the remaining Al" phase forms a eutectoid structure of a fine Al phase and a fine Zn phase at temperature corresponding to the eutectoid reaction in

the Al-Zn binary equilibrium state diagram.

Fig. 2 (a) is a photomicrograph (magnification: 3,500 x) that enlarges the Al" phase in Fig. 1. Fig. 2 (b) is a view showing the distribution state of each structure in the photomicrograph. It can be concluded from the observation in detail of the Al" phase that eutectoid structures that are condensed precipitated Zn phases are present outside the Al" phase and around the Ti-Al base intermetallic compound.

Although there is no specific restriction on the size of the intermetallic compound in the present invention, the size observed by the present inventors is up to 10 μm . Moreover, there is no specific limitation on the proportion of the intermetallic compound present in the plating layer structure. However, it is desirable that the intermetallic compound be present in an amount of at least 10% in one of the [Al phase], [Zn₂Mg phase] and [Zn phase].

There is no specific restriction on the method of adding the intermetallic compound. A method of dispersing fine powder of the intermetallic compound in the bath, a method of dissolving the intermetallic compound in the bath, and the like method, can be applied. When the plated steel product is produced in a continuous line in which hot-dip galvanizing with a nonoxidizing furnace system is used, a method of dissolving Ti in the plating bath is suitable. As the method of dissolving Ti in a plating bath, a method of adding a Ti-Zn base intermetallic compound is efficient because Ti can be dissolved at low temperature in a short period of time. Examples of the Ti-Zn base intermetallic compound to be added include Zn₁₅Ti, Zn₁₀Ti, Zn₅Ti, Zn₃Ti, Zn₂Ti and ZnTi. When such an intermetallic compound is added to the plating bath singly or in a mixture of the intermetallic compound and Zn, or an alloy of Zn-Al or Zn-Al-Mg, dissolved Ti crystallizes in the plating layer as a Ti-Al base intermetallic compound, and the compound

improves the surface smoothness and formability.

Examples of the substrate steel product of the present invention include not only steel sheets but also various steel products such as wire rods, shape steels, bar steels and steel tubes. Usable steel sheets include both hot rolled steel sheets and cold rolled steel sheets. Various steel types such as Al-killed steels, extra low carbon steels containing Ti, Nb, etc., high-strength steels containing strengthening elements such as P, Si and Mn and stainless steels can be used.

There is no specific restriction on the production process of the steel products in the invention, and various processes such as continuous plating of steel sheets, and hot-dip galvanizing of steel products and wire rods can be applied. When steel products are to be pre-plated with Ni as a substrate layer, a conventional pre-plating process may be applied. Because the plated products obtained in the present invention each have a plating layer excellent in surface smoothness even when the cooling rate is small, the process shows a significant effect in hot-dip galvanizing in which the plated products are hardly cooled at a large cooling rate, and hot-dip galvanizing of thick steel products.

Although there is no specific limitation on the adhesion amount of the plating layer, the amount is desirably at least 10 g/m² in view of the corrosion resistance, and up to 350 g/m² in view of the workability.

The zinc plating layer may also contain Fe, Sb, Pb and Sn singly or in a mixture in an amount of 0.5% by mass or less in addition to the above elements. Moreover, even when the plating layer contains Ca, Be, Cu, Ni, Co, Cr, Mn, P, B, Nb, Bi and group III elements in a total amount of 0.5% by mass or less, the effect of the invention is not impaired, and the corrosion resistance of the plated steel products is sometimes further preferably improved, depending on the amount.

Examples

(Example 1)

First, cold rolled steel sheets 0.85 mm thick were prepared. Each steel sheet was hot-dip galvanized for 3
5 sec in a plating bath at temperature of 400 to 600°C. Amounts of addition elements in the bath were varied. The adhesion amount of a plating layer on one side of each steel sheet was adjusted to 140 g/m² by N₂ wiping, and the plated steel sheet was cooled at a rate of
10 10°C/sec or less. Table 1 shows the plating compositions of the plated steel sheets thus obtained. The cross section of each steel sheet was observed with a SEM, and Table 1 also shows the results of observing the metal structure of the plating layer.

15 Each plated steel sheet was inclined at an angle of 10° to the horizontal plane, and surface ground. Ti-Al base intermetallic compounds present in an [Al phase], a [Zn₂Mg phase] and a [Zn phase] were subsequently observed with an EPMA.

20 The size of dendrites in [Al phases] in the plating layer was determined by the following procedure. The surface of each plated steel sheet was mapped by CMA, and major axes of dendrites in the Al map thus obtained were measured in an area of 5 x 5 cm. The major axes of five
25 dendrites were measured in order of decreasing magnitude. The average was used as the size of dendrites in the [Al phases].

As to the smoothness, R_a and W_{CA} were measured with a surface roughness shape measurement apparatus
30 (manufactured by Tokyo Seimitsu Co., Ltd.). The surface roughness of each steel sheet (only solidified by cooling) was measured at optional 3 sites under the conditions explained below, and the average was used.

Measurement probe: stylus tip having a curvature of

35 5 μm R

Measurement length: 25 mm

Cut off: R_a 0.8 mm, W_{CA} 0.8 to 8 mm

Driving speed: 0.3 mm/sec

Filter: 2 CR filter

5 The smoothness of each steel sheet was judged from the following scores. A steel sheet having a score of 4 was accepted.

4: R_a up to 1 μm , W_{CA} up to 1 μm

3: R_a exceeding 1 μm , W_{CA} up to 1 μm

2: R_a up to 1 μm , W_{CA} exceeding 1 μm

10 1: R_a exceeding 1 μm , W_{CA} exceeding 1 μm

The formability of each steel sheet was evaluated by a draw bead test. A drawing load obtained under the following measurement conditions was used, and the apparent friction coefficient was calculated.

15 Bead mold: round shape of a projected portion R 4 mm R, shoulder R 2 mm R

Sample size: 30 mm x 300 mm

Slide length: 110 mm

Drawing speed: 200 mm/min

20 Pressing load: 600, 800, 1,000 kgf

The smoothness of each steel sheet was judged from the following scores. A steel sheet having a score of 3 was accepted.

3: less than 0.20

25 2: at least 0.20 to less than 0.21

1: at least 0.21

Each steel sheet was sprayed with 5% salt water at 35°C for 1,000 hours. When rust was not formed, the steel sheet was accepted. When rust was formed, the steel sheet was rejected.

30 Table 1 shows the evaluation results. Because Sample No. 14 contained no Ti-Al base intermetallic compound, an Al phase grew, and it was rejected because of the smoothness and formability. Because Sample No. 15 had an excessive Ti content, a Ti-Al base intermetallic compound condensed in the surface, and it was rejected

because of the smoothness and formability. Because
Sample No. 16 had contents of Mg, Al, Si and Ti outside
the scope of the present invention, it was rejected
because of the corrosion resistance. Samples other than
5 the above ones each showed good smoothness, formability
and corrosion resistance.

Table 1

Sample No.	Composition of hot-dip Zn coating layer (mass%)				Metal structure					
	Mg	Al	Si	Ti	Mg ₂ Si phase	Ternary eutectic	Al phase	Zn phase	MgZn ₂ phase	Ti-Al base inter-metallic cpd.
1	4	8	0.15	0.009	○	○	○	○	○	○
2	5	10	0.2	0.009	○	○	○		○	○
3	5	15	0.45	0.009	○	○	○		○	○
4	3	6	0.05	0.009	○	○	○	○		○
5	1	19	0.5	0.009	○	○	○		○	○
6	1	4	0.005	0.009	○	○	○	○		○
7	3	11	0.2	0.009	○	○	○		○	○
8	3	11	0.2	0.001	○	○	○		○	○
9	3	11	0.2	0.0002	○	○	○		○	○
10	3	11	0.2	0.00005	○	○	○		○	○
11	3	11	0.2	0.000001	○	○	○		○	○
12	3	11	0.0002	0.009	○	○	○		○	○
13	3	11	0.000001	0.009	○	○	○		○	○
14	3	11	0.2	0	○	○	○		○	○
15	3	11	0.2	0.12	○	○	○		○	○
16	0	0.2	0	0				○		
17	3	6	0	0.01		○	○	○	○	○

5

Table 1 (Continued)

Sample No.	Size of Al phase	Smoothness	Formability	Corrosion resistance	Note
1	150 μm	4	3	Accepted	Ex.
2	150 μm	4	3	Accepted	Ex.
3	150 μm	4	3	Accepted	Ex.
4	100 μm	4	3	Accepted	Ex.
5	200 μm	4	3	Accepted	Ex.
6	100 μm	4	3	Accepted	Ex.
7	150 μm	4	3	Accepted	Ex.
8	200 μm	4	3	Accepted	Ex.
9	300 μm	4	3	Accepted	Ex.
10	350 μm	4	3	Accepted	Ex.
11	450 μm	4	3	Accepted	Ex.
12	150 μm	4	3	Accepted	Ex.
13	150 μm	4	3	Accepted	Ex.
14	1200 μm	2	2	Accepted	Comp.Ex.
15	200 μm	3	2	Accepted	Comp.Ex.
16	-	4	3	Rejected	Comp.Ex.
17	150 μm	4	3	Accepted	Ex.

(Example 2)

First, cold rolled steel sheets 0.85 mm thick were prepared. Each steel sheet was hot-dip galvanized for 3 sec in a plating bath at a temperature of 520°C. Amounts of addition elements in the bath were varied. The
5 adhesion amount of a plating layer on one side of each steel sheet was adjusted to 140 g/m² by N₂ wiping, and the plated steel sheet was cooled at a rate of 10°C/sec or less. Table 2 shows the plating compositions of the
10 plated steel sheets thus obtained. The cross section of each steel sheet was observed with a SEM, and Table 2 also shows the results of observing the metal structure of the plating layer.

Each plated steel sheet was inclined at an angle of
15 10° to the horizontal plane, and surface ground. Ti-Al base intermetallic compounds present in an [Al phase], a [Zn₂Mg phase] and a [Zn phase] were subsequently observed with an EPMA. Moreover, as to Ti-Al base intermetallic compounds present in an [Al phase], the presence or
20 absence of the Ti-Al base intermetallic compounds in a Zn-Al eutectoid structure in which a precipitated [Zn phase] was condensed was observed with an EPMA. Furthermore, whether a Ti-Al base intermetallic compound contained Si or not was observed by observing a Ti-Al
25 base intermetallic compound with an EPMA.

The adhesion of the plating layer was evaluated by the following procedure. An adhesive cellophane tape was affixed to each hot-dip galvanized steel sheet subsequent to a Du Pont impact test, and the tape was peeled off the
30 steel sheet. The adhesion was evaluated by the following criteria:

O: the plating layer is not exfoliated;

Δ: the plating layer is exfoliated in an area of less than 10%; and

35 X: the plating layer is exfoliated in an area of at least 10%. In the Du Pont test, an impact mold having a

roundness of 1/2 inch at the tip was used, and the test was carried out by dropping a 1-kg weight from a height of 1 m.

5 The corrosion resistance of each plated steel sheet subsequent to working was evaluated by the following procedure. A sample was subjected to 1 T bending (bending a sample to be tested at an angle of 180° while the sample was being held). The bent portion of the sample was subjected to SST for 1,000 hours. The rust
10 formation state on the bent portion was judged according to the following scores. A steel sheet having a score of at least 3 was accepted.

- 5: rust formation area of less than 5%;
- 15 4: rust formation area of from at least 5% to less than 10%;
- 3: rust formation area of from at least 10% to less than 20%;
- 2: rust formation area of from at least 20% to less than 30%; and
- 20 1: rust formation area of at least 30%.

Table 2 shows the evaluation results. Because the addition amounts of Al and Si in Sample No. 2 were outside the scope of the present invention, Sample No. 2 was rejected because of the adhesion. The other samples
25 each showed good adhesion of the plating layer and good corrosion resistance after working. Plated steel sheets that had a plating layer containing Si showed particularly good adhesion and corrosion resistance after
30 working.

Table 2

Sample No.	Composition of hot-dip Zn plating layer (mass%)				Metal structure					
	Mg	Al	Si	Ti	Mg ₂ Si Phase	Ternary eutectic	Al phase	Zn phase	MgZn ₂ phase	Ti-Al base inter-metallic cpd.
1	3	12	0.2	0.009	○	○	○		○	Ti(Al _{0.85} Si _{0.15}) ₃
2	3	12	0	0.009		○	○		○	TiAl ₃
3	3	6	0	0.01		○	○		○	TiAl ₃
4	3	6	0.002	0.01	○	○	○		○	Ti(Al _{0.85} Si _{0.15}) ₃

5 Table 2 (Continued)

Sample No.	Site where Ti-Al base intermetallic compound in Al phase was present	Adhesion	Corrosion resistance after working	Note
1	In Zn-Al eutectoid structure where Zn phase was condensed	○	5	Ex.
2	In Zn-Al eutectoid structure where Zn phase was condensed	X	3	Comp.Ex.
3	In Zn-Al eutectoid structure where Zn phase was condensed	△	4	Ex.
4	In Zn-Al eutectoid structure where Zn phase was condensed	○	5	Ex.

Industrial Applicability

As explained above, according to the present invention, highly corrosion-resistant plated steel products that show excellent surface smoothness and formability, without forming protrusions and recesses in the surfaces even when the solidification rate of the plating layers was small, can be produced.